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Quantum Numbers, Atomic Orbitals, and Electron Configurations 2. Atomic Structure LECTURE 1 CHAPTER 2 : ATOMIC STRUCTURE ( Video 1 )

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## Topic 2: Atomic Structure | ib-chemistry

2.1 The nuclear atom. Atom Atoms consist of a three subatomic particles Proton Neutron Electron Atomic Number Z: number of protons in the nucleus of an atom of an element. Mass Number A: number of protons + number of neutrons Relative atomic mass  $A_r$ : Ratio of the average mass of the atom to the  $^{12}\text{C}$ . Continue reading. 2.1 The nuclear atom.

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months ago. ncamarena. 0. Save. Edit. Edit. ... The average atomic mass rhenium, Re, is 186.21 amu. If 37.1% of rhenium has mass# = 185, what is the other stable isotope? answer choices

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Atomic Structure. 2.1 The Nuclear atom. 2.1.1 : Protons -- Mass = 1 amu , charge = +1 .. Neutrons -- Mass = 1 amu , charge = 0 .. Electron -- Mass = 1/1840 amu (usually insignificant), charge = -1. 2.1.2 : Protons and Neutrons form the nucleus of the atom, electrons orbit the nucleus in electron shells.

2 Atomic Structure - EARJ IB CHEMISTRY

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A number that is equal to the number of protons in the nucleus of an atom.

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Atoms contain a positively charged dense nucleus composed of protons and neutrons (nucleons). Negatively charged electrons occupy the space outside the nucleus. The main energy level or shell is given an integer number,  $n$ , and can

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hold a maximum number of electrons,  $2n^2$ . A more detailed model of the atom describes the.

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We call this second type of orbital a 'p' orbital. Carbon holds no surprises and the expected value (6/5 times that of boron) is  $1086 \text{ kJ mol}^{-1}$ . The slightly large value can be ascribed to 'shrinkage' once again. Nitrogen also follows the pattern as its first ionisation energy is  $1420 \text{ kJ mol}^{-1}$ .



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species W, X, Y and Z below. Which species is an anion? A.  
W. B. X.

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chemically combining two or more elements. 1 of 11. Atomic  
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Atomic Structure Atomos is an ancient Greek word meaning 'indivisible' or 'cannot be cut'. Atoms are the building blocks of matter; they are the smallest particles that when combined together have the properties of a particular chemical element, or compound.

StudyIB Chemistry: Atomic Structure

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